

St. FRANCIS COLLEGE FOR WOMEN, BEGUMPET HYDERABAD-500016
(An Autonomous College Affiliated to Osmania University)
DEPARTMENT OF CHEMISTRY

DISCIPLINE SPECIFIC CORE I
CHEMISTRY I
B.Sc. I - SEMESTER- I 60 h

Module 1: Inorganic Chemistry

Chemistry of p-Block elements

Module 2: Organic Chemistry

Structural Theory in Organic Chemistry

Acyclic Hydrocarbons

Aromatic Hydrocarbons

Module 3: Physical Chemistry

Elementary Quantum Mechanics

Chemical Kinetics

Photochemistry

Module 4: General Chemistry

General Principles of Inorganic Quantitative Analysis

Isomerism

Colloids & Surface Chemistry

B. Saritha

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**SEMESTER - I
CHEMISTRY - I****1. Course Description**

Programme: B.Sc.
Course Code: U26/CHE/DSC/101
Course type: DSC - 1
No. of credits: 4

Max. Hours: 60
Max. Marks: 100
Hours per week: 4

2. Course Objectives

- To develop an understanding of the periodic trends, electronic configuration, bonding, structure, and reactivity of p-block elements and their compounds.
- To enable students to comprehend structural theory, bonding, stereochemistry, and the preparation, properties, and reactions of acyclic and aromatic hydrocarbons.
- To introduce the fundamental principles of quantum mechanics, chemical kinetics, and photochemistry, and apply them to explain molecular behavior and reaction mechanisms.
- To equip students with foundational knowledge of inorganic quantitative analysis, structural and stereoisomerism, and the principles governing colloids and surface phenomena.

3. Course Outcomes

On completion of the course the student will be able to

- Analyze the electronic configuration, periodic trends, bonding, and reactivity of p-block elements to predict the properties of their compounds.
- Apply principles of structural theory and mechanistic reasoning to interpret and predict the reactions of acyclic and aromatic hydrocarbons.
- Explain and apply the fundamental concepts of quantum mechanics, chemical kinetics, and photochemistry to analyze molecular systems and reaction rates.
- Demonstrate analytical and conceptual understanding of inorganic quantitative analysis, isomerism, and colloidal and surface phenomena in chemical systems.

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Alkynes— Preparation by dehydrohalogenation of vicinal dihalides, dehalogenation of tetrahalides. Physical Properties: Chemical reactivity – electrophilic addition of X_2 , HX , H_2O (tautomerism), Oxidation (formation of enediol) and reduction (catalytic hydrogenation).

Aromatic Hydrocarbons

5h

Introduction to aromaticity: Huckel's rule – Benzene, Naphthalene and Anthracene. Reactions - General mechanism of electrophilic substitution, mechanism of nitration, sulphonation and halogenation, Friedel Crafts alkylation and acylation. Orientation of aromatic substitution - Definition of ortho, para and meta directing groups. Ring activating and deactivating groups with examples. Orientation – (i) activating groups: Amino, methoxy and alkyl groups. (ii) Deactivating groups - nitro, nitrile, carbonyl, carboxylic acid, sulphonic acid and halo groups.

IKS: Traditional distillation methods in the extraction of perfumes.

MODULE III: PHYSICAL CHEMISTRY

15 h (1 h/week)

Elementary quantum mechanics

3h

Limitations of classical mechanics and Origin of quantum mechanics-Black body radiation, Rayleigh Jeans law; Planck's radiation law, photoelectric effect, Compton effect, de Broglie's hypothesis. Heisenberg's uncertainty principle. Schrödinger wave equation (derivation not required) – significance of ψ and ψ^2 .

Chemical Kinetics

8h

Introduction to chemical kinetics, rate of reaction, rate laws and rate constant. Molecularity and Order of a reaction. Factors influencing the reaction rates. First order reaction, derivation of equation for rate constant. Characteristics of first order reaction. Units for rate constant. Half- life period, graph of first order reaction, Example - Decomposition of H_2O_2 . Problems. Pseudo first order reaction, Hydrolysis of methyl acetate, inversion of cane sugar, problems. Second order reaction, derivation of expression for second order rate constant, example-Saponification of ester. Characteristics of second order reaction, units for rate constants, half- life period and second order plots. Problems. Methods for determining the order of a reaction. Arrhenius equation – activation energy -problems.

Photochemistry

4h

Introduction to photochemistry – differences between dark and photo reactions. Laws of photochemistry; Quantum Yield – problems; Examples of photo chemical reactions with different quantum yields. Photo chemical combinations of H_2-Cl_2 and H_2-Br_2 reactions. Abnormal quantum yield – high and low-examples with reasons. Singlet and triplet states. Jablonski diagram – non radiative processes – Internal conversion and Intersystem crossing; radiative processes- Fluorescence and phosphorescence.

4. Course Content

MODULE I: INORGANIC CHEMISTRY

15 h (1 h/week)

Chemistry of p-Block Elements

15 h

Structure and bonding in diborane (B_2H_6), Boron nitrogen compounds ($B_3N_3H_6$ and BN), Lewis acid nature of BX_3 .

Carbides- Classification -ionic, covalent, interstitial-Structures and reactivity. Industrial applications. Silicones-Classification-straight chain, cyclic and cross-linked and applications.

Nitrides-Classification -ionic, covalent and interstitial- Reactivity – hydrolysis.

Oxides and Oxyacids- Definition and Types of oxides (a) Normal- acidic, basic amphoteric and neutral (b) Mixed oxide (c) sub oxide (d) peroxide (e) superoxide. Structure of oxides and oxyacids of B, C, N, P, S and Cl - reactivity, thermal stability, hydrolysis.

Interhalogens- Classification- general preparation- structures of AB , AB_3 , AB_5 and AB_7 type and reactivity.

Poly halide- Definition and structure of ICl_2^- , ICl_4^- and I_3^- .

Pseudohalogens- Comparison with halogens.

Structure, bonding and reactivity of Xenon Compounds-Oxides, Halides and Oxy-halides.

MODULE II: ORGANIC CHEMISTRY

15 h (1 h/week)

Structural Theory in Organic Chemistry

5h

Bond polarization: Factors influencing the polarization of covalent bonds, electro negativity – inductive effect. Application of inductive effect (a) Basicity of amines (b) Acidity of carboxylic acids (c) Stability of carbonium ions. Resonance - Mesomeric effect, application to (a) acidity of phenol. (b) acidity of carboxylic acids and basicity of anilines. Stability of carbo cations, carbanions and free radicals. Hyper conjugation and its application to stability of carbonium ions, free radicals and alkenes.

Acyclic Hydrocarbons

5h

Alkanes– Methods of preparation: Preparation of Alkanes from Grignard reagent. Chemical reactivity- inert nature, free radical substitution, Halogenation example.

Alkenes - Preparation of alkenes (with mechanism) (a) by dehydration of alcohols (b) dehydrohalogenation of alkyl halides (c) by dehalogenation of 1,2 dihalides, Zaitsev's rule. Properties: Anti-addition of halogen and its mechanism. Addition of HX, Markovnikov's rule, addition of H_2O , HOX with mechanism and addition of HBr in the presence of peroxide (anti Markovnikov's addition). Oxidation (cis-additions) hydroxylation by $KMnO_4$, OsO_4 , anti-addition- peracids (via epoxidation), ozonolysis – location of double bond.

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MODULE IV: GENERAL CHEMISTRY

15 h (1 h/week)

General Principles of Inorganic Quantitative Analysis 5h

Volumetric Analysis: Introduction, standard solutions, indicators, end point, titration curves, Types of titrations: i) neutralization titration- principle, theory of acid base indicators, titration curves and selection of indicators- strong acid - strong base, strong acid -weak base, weak acid- strong base and weak acid -weak base. Theory of redox titrations – internal (KMnO_4) and external indicators – use of diphenylamine and ferroin indicators. Theory of complexometric titrations – use of EBT, Murexide and Fast sulphone black indicators. Role of pH in complexometric titrations. Precipitation titrations – theory of adsorption indicators.

Isomerism 5h

Isomerism: Definition of isomers. Classification of isomers: Constitutional and Stereoisomers - definition and examples. Constitutional isomers: chain, functional and positional isomers. Stereoisomers: enantiomers and diastereomers – definitions and examples. Representation of stereoisomers – Wedge, Fischer, Sawhorse, Newmann projection formulae.

Conformational analysis: Classification of stereoisomers based on energy. Definition and examples Conformational and configurational isomers. Conformational analysis of ethane, n- butane, 1,2- dichloroethane, 2-chloroethanol. Cis-trans isomerism: E-Z Nomenclature.

Colloids & Surface Chemistry 5h

Colloids: Definition of colloids-classification of colloids-examples. Solid in liquid (sol)-Preparation, kinetic and electrical properties, stability and protection of colloids - Hardy-Schulze rule and Gold number. Liquid in liquid (emulsion)-types of emulsions and emulsifiers. Liquid in solid (gel)-types and properties. Applications of colloids. **Adsorption:** Types of adsorptions; Factors influencing adsorption; Freundlich adsorption isotherm and Langmuir adsorption isotherm. Applications.

5. Reference Books

General reference: B.Sc I Year Chemistry : Semester I, Telugu Academy publication, Hyd.

Module- I

1. Puri, B. R., Sharma, L. R., & Kalia, M. S. (1996). Principles of inorganic chemistry. Vishal Publications.
2. Lee, J. D. (1981). Concise inorganic chemistry (3rd ed.). Oxford University Press.
3. Cotton, F. A., Wilkinson, G., & Gaus, P. L. (2001). Basic inorganic chemistry (3rd ed.). Wiley.

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- Huheey, J. E., Keiter, E. A., & Keiter, R. L. (1993). Inorganic chemistry: Principles of structure and reactivity (4th ed.). Harper Collins College Publishers.
- Greenwood, N. N., & Earnshaw, A. (1989). Chemistry of the elements. Pergamon Press.
- Shriver, D. F., & Atkins, P. W. (1999). Inorganic chemistry (3rd ed.). Oxford University Press.
- Gopalan, R. (2009). Textbook of inorganic chemistry. University Press.

Module- II

- Morrison, R. T., & Boyd, R. N. (2011). Organic chemistry. Pearson Education (Prentice Hall).
- Solomons, T. W. G., & Fryhle, C. B. (2016). Organic chemistry. Wiley (John Wiley & Sons).
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- Soni, P. L., & Soni, H. M. (2012). Organic chemistry. Sultan Chand & Sons.
- Ghosh, S. K. (2009). General organic chemistry. Bharati Bhawan Publishers.
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Module- III

- Puri, B. R., Sharma, L. R., & Pathania, M. S. (2013). Principles of physical chemistry (46th ed.). Vishal Publishing Company.
- Raj, G. (2009). Advanced physical chemistry (35th ed.). Goel Publishing House.
- Lewis, G., & Glasstone, S. (1966). Elements of physical chemistry. Macmillan.
- Atkins, P. W. (2001). Physical chemistry (7th ed.). Oxford University Press.
- Kapoor, K. L. (1994). A textbook of physical chemistry (Vols. 4 & 5). Macmillan India Ltd.
- Laidler, K. J. (1987). Chemical kinetics (3rd ed.). McGraw Hill.
- Rajaraman, J., & Kuriacose, J. (1993). Kinetics and mechanism of chemical transformations. Macmillan India.
- Turro, N. J. (1978). Molecular photochemistry. W. A. Benjamin, Inc.
- Rohatgi-Mukherjee, K. K. (1978). Fundamentals of photochemistry. Wiley Eastern.
- Dogra, S. K., & Dogra, S. (1996). Physical chemistry through problems (4th ed.). New Age International.
- Kalidas, C., & Sangaranarayanan, M. V. (2019). Physical chemistry: Problems and solutions. University's Press.

Module- IV

- Jeffery, G. H., Bassett, J., Mendham, J., & Denney, R. C. (1999). Vogel's textbook of quantitative chemical analysis (5th ed.). Addison Wesley Longman Inc.
- Day, R. A., & Underwood, A. L. (2004). Quantitative analysis (6th ed.). Prentice Hall of India.

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3. Svehla, G. (1996). Vogel's qualitative inorganic analysis (7th ed.). Prentice Hall.
4. Morrison, R. T., & Boyd, R. N. (2011). Organic chemistry. Pearson Education.
5. Solomons, T. W. G., & Fryhle, C. B. (2016). Organic chemistry. Wiley.
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10. Atkins, P., & de Paula, J. (2010). Atkins' physical chemistry (9th ed.). Oxford University Press.
11. McQuarrie, D. A., & Simon, J. D. (1997). Physical chemistry: A molecular approach. Viva Books Pvt. Ltd.
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6. Syllabus Focus

a. Relevance to Local, Regional, National and Global Development Needs

Local/Regional/ National/Global Development Needs	Relevance
Local	Helps individuals understand and safely use everyday chemical products such as fuels, medicines, fertilizers, and cleaning agents.
Regional	Supports regional industries like agriculture, petrochemicals, pharmaceuticals, and environmental monitoring.
National	Contributes to national economic growth through industrial development, technological innovation, and scientific research.
Global	Plays a vital role in addressing global challenges such as climate change, sustainable energy, food security, and healthcare advancements.

b. Components on Skill Development/Entrepreneurship Development/Employability

SD/ED /EMP	Syllabus Content	Description of Activity

SD	Module I	Encourage them to compare and discuss trends in reactivity and chemical behaviour.
EMP	Module II	Students will create a brief dossier on a hydrocarbon compound, highlighting its structure–property link, synthesis, safety, and industrial uses to build employability skills.
ED	Module III	Students will develop a startup proposal for a light-driven or catalytic process applying concepts of quantum mechanics, chemical kinetics, and photochemistry to assess feasibility and market potential.
EMP NSQF level 3 & 4	Module IV	Hands on practical training and field visits.

c. IKS component:

IKS	Syllabus Content	Module
	Introduced Traditional distillation methods in the extraction of perfumes.	II

7. Pedagogy

S. No.	Student Centric Methods Adopted	Type / Description of Activity
1	Experiential Learning	Micro-scale experiments in Volumetric analysis and field visits.
2	Participative Learning	Collage/ Quiz/ JAM/Game based

8. Course Assessment Plan

a) Weightage of Marks in Continuous Internal Assessments and End Semester Examination

CO	Continuous Internal Assessments CIA - 40%	End Semester Examination-60%
CO1	CIA-1: Written Exam	Written Exam

CO2	CIA-2: Collage/Quiz/JAM/	
CO3	CIA-1: Written Exam	
CO4	CIA-2: Assignment/ Game based/ Poster making	

b. Model Question Paper - End Semester Exam Theory

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 (An Autonomous College Affiliated to Osmania University)
 Faculty of Science – Department of Chemistry
MODEL PAPER
B.Sc. I YEAR SEMESTER - I
CHEMISTRY - I

Time: 2 hrs

Course Code: U26/DSC/CHE/101

Max. Marks: 60

SECTION –A (Essay Questions)

I. Answer the following

4x 10=40 Marks

- Classify the oxides based on the oxygen content. 4M
 - What are interhalogen compounds? Explain the structure of AX₃ type of molecules. 6M

OR
- Discuss the structure of Diborane. 5M
 - What are silicones? Give their classification. 5M
- What is the Mesomeric effect? How does it explain the acidity of phenols? 5M
 - Explain the acidic strength of carboxylic acids by applying the concept of Inductive effect. 5M

OR
- Write the mechanism for Friedel Crafts alkylation of benzene. 5M
 - What are ortho, meta and para directing groups? Explain with one example each. 5M
- Write the Schrodinger wave equation and explain the significance of ψ and ψ^2 and draw shapes of p and d atomic orbitals. 10M

OR
- Derive an expression for the rate constant of first order reaction. 5M

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- b. A first order reaction is 50% complete in 100 minutes. How long will it take for 90% completion? 5M
7. Explain the principle involved in Redox titrations? Add a note on the detection of end points. 10M
- OR
8. a. Classify various types of colloids. 5M
- b. Deduce the expression for Langmuir adsorption isotherms. 5M

SECTION –B (Short Answer Questions)

II. Answer any four.

4x5=20 Marks

9. Discuss the structure of XeO₃.
10. State and explain Markovnikov's rule with mechanism.
11. What is Huckel's rule? State whether naphthalene, cyclopentadienyl anion and tropylium cation are aromatic or not based on Huckel's rule.
12. Determine the order of reaction using Van't Hoff differential method.
13. State and explain Grothus and Draper's law.
14. Discuss the conformations of n-Hexane. Which is more stable and why?

Question Paper Format - Blooms Taxonomy Level

SECTION A - INTERNAL CHOICE				4 X 10 = 40 M	
Question Number	Module	Question	CO	BTL	
1	I	a. Classify the oxides based on the oxygen content. 4M b. What are interhalogen compounds? Explain the structure of AX ₃ type of molecules. 6M OR	CO1	Level 1 & 2	
2	I	a. Discuss the structure of Diborane. (CO1) 5M b. What are silicones? Give their classification. 5M	CO1	Level 1 & 2	

3	II	a. What is the Mesomeric effect? How does it explain the acidity of phenols? 5M b. Explain the acidic strength of carboxylic acids by applying the concept of Inductive effect. 5M OR	CO2	Level 3
4	II	a. Write the mechanism for Friedel Crafts alkylation of benzene. 5M b. What are ortho, meta and para directing groups? Explain with one example each. 5M	CO2	Level 2 & 3
5	III	a. Write the Schrodinger wave equation and explain the significance of ψ and ψ^2 and draw shapes of p and d atomic orbitals. 10M OR	CO3	Level 2
6	III	a. Derive an expression for the rate constant of first order reaction. 5M b. A first order reaction is 50% complete in 100 minutes. How long will it take for 90% completion? 5M	CO3	Level 4 & 5
7	IV	Explain the principle involved in Redox titrations? Add a note on the detection of end points. 10M OR	CO4	Level 2
8	IV	a. Classify various types of colloids. 5M b. Deduce the expression for Langmuir adsorption isotherms. 5M	CO4	Level 2 & 4
SECTION B - Short answer questions				
ANSWER ANY 4 OUT OF 6				4 x 5 = 20M
9	I	Discuss the structure of XeO_3 .	CO1	Level 2
10	II	State and explain Markovnikov's rule with mechanism.	CO2	Level 1 & 4
11	II	What is Huckel's rule? State whether naphthalene, cyclopentadienyl anion	CO2	Level 3

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		and tropylium cation are aromatic or not based on Huckel's rule.		
12	III	Determine the order of reaction using Van't Hoff differential method.	CO3	Level 3
13	III	State and explain Grothus and Draper's law.	CO3	Level 2
14	IV	Discuss the conformations of n-Hexane. Which is more stable and why?	CO4	Level 2 & 4

c) Question Paper Blueprint

Modules	Hours Allotted in the Syllabus	COs Addressed	Section A (No. of Questions)	Total Marks	Section B (No. of Questions)	Total Marks
1	15	CO1	2	10	6 (By taking at least one question from each module)	4x5=20
2	15	CO2	2	10		
3	15	CO3	2	10		
4	15	CO4	2	10		

9. CO-PO Mapping

CO	PO	Cognitive Level	Classroom sessions (h)
1	1,3	Understand	15
2	1,2	Analyse	15
3	1,2	Apply	15
4	1,7	Remember	15

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FACULTY OF SCIENCE- DEPARTMENT OF CHEMISTRY
SEMESTER -I
LABORATORY COURSE-I - QUANTITATIVE ANALYSIS

1. Course description

Program: B.Sc.

Max. Hours: 20

Course Code: U26/CHE/DSC/101/ P

Max. Marks: 50

Course type: DSC-1

Hours per week: 2

No. of Credits: 1

2. Course Objective

- To enable students to understand the principles of volumetric analysis and apply titrimetric methods for precise quantitative estimation of chemical substances.

3. Course outcomes

CO 1: Students will be able to accurately prepare and standardize solutions, select suitable indicators, and perform various titrimetric methods with proper laboratory technique.

CO 2: Students will be able to apply volumetric analysis techniques to determine the composition or strength of constituents in market samples and interpret the results with appropriate calculations.

4. Course content**QUANTITATIVE ANALYSIS****Acid-Base Titrations**

- Estimation of Carbonate in Washing Soda.
- Estimation of Bicarbonate in Baking Soda.
- Estimation of Carbonate and Bicarbonate in the Mixture.
- Estimation of Alkali content in Antacid using HCl.

Redox Titrations

- Determination of Fe(II) using $K_2Cr_2O_7$
- Determination of Fe(II) using $KMnO_4$ with sodium oxalate as primary standard.
- Determination of Cu(II) using $Na_2S_2O_3$ with $K_2Cr_2O_7$ as primary standard

Complexometric Titrations

- Estimation of Mg^{2+} by EDTA
- Estimation of Cu^{2+} by EDTA

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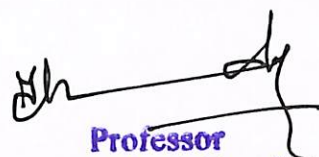
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5. Reference Books

1. Jeffery, G. H., Bassett, J., Mendham, J., & Denney, R. C. (1999). Vogel's textbook of quantitative chemical analysis (5th ed.). Addison Wesley Longman Inc.
2. Vogel, A. I. Vogel's Textbook of Quantitative Chemical Analysis, 6th Edition, Pearson Education, 2000. ISBN: 9780582226289
3. Giri, A. N. (2010). A textbook of practical chemistry. Himalaya Publishing House.
4. O.P. Pandey, D.N. Bajpai, & S. Giri. (2020). Practical chemistry. 10th Revised Edition, S. Chand Publishing. ISBN: 9789352535859
5. Gopalan, R., Subramanian, P. S., & Raghavan, K. (2004). Elements of analytical chemistry. Sultan Chand & Sons.
6. Gopalan, R., Venkappayya, D., and Nagarajan, S. (2012). Textbook of Inorganic Chemistry (Lab Manual), 3rd Edition, Universities Press, Hyderabad, ISBN: 9788173718204
7. Ahluwalia, V. K., and Sunita Dhingra, (2005). A Laboratory Manual of Organic and Inorganic Chemistry, 1st Edition, University Press, Hyderabad, ISBN: 9788173715623


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6. Model Question Paper - End Semester Exam Practical

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(An Autonomous College Affiliated to Osmania University)

Faculty of Science – Department of Chemistry

MODEL PAPER

B.Sc. I YEAR SEMESTER -I

LABORATORY COURSE-I - QUANTITATIVE ANALYSIS

Program: B.Sc.

Course Code: U26/DSC/CHE/101/P

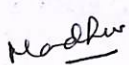
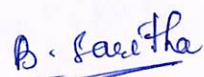

Type of course: DSC-1

Time: 2 hrs

No. of credits: 1

Max. Marks: 50


1. Estimate the amount of Oxalic acid present in 100 mL of the given unknown solution. You are provided with a pure sample of sodium oxalate and KMnO_4 solution of 0.02 M approximately. 30M
2. Viva 10 M
3. Record & attendance 10 M

Prepared by	Checked & Verified by	Approved by
 Dr Lakshmi Madhuri Ms B Prashanthi	 Dr Saritha Aduri HoD	 Prof. Uma Joseph Principal Principal

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SEMESTER I
Core Course 1
CHEMISTRY OF BIOMOLECULES

Code : U26/BIC/DSC/101

Credits : 4

Total hours : 60

Hours/week : 4

Module I : WATER & BUFFERS

Module II : CARBOHYDRATES

Module III : AMINO ACIDS & PROTEINS

Module IV : LIPIDS

CHEMISTRY OF BIOMOLECULES

1. Course Description

Programme : B.Sc.
Course Code : U26/BIC/DSC/101
Type of course : DSC 1
No. of credits : 4

Max. Hours : 60
Hours per week: 4
Max. Marks: 100

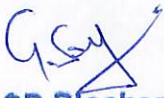
2. Course Objectives


1. To understand the basic molecules of life about the structures, properties, functions and classification.
2. To elucidate the importance of knowledge on the biochemical aspects of the biomolecules in biological systems.

3. Course Outcomes

After the successful completion of the course, the student will be able to:

- CO1: Understand basic biochemical principles related to water, acids and bases, pH and buffer systems, and functional groups in biomolecules. (L1)
- CO2: Understand the classification, structure, and biological roles of carbohydrates. (L1)
- CO3: Summarize the structures of amino acids, reactions of amino acids, protein classification, different levels of organization of proteins(L2)
- CO 4: Illustrate the structure and function of lipids, fatty acids and lipoproteins(L3)


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4. Course Content:

MODULE 1: WATER & BUFFERS

15 Hours

Scope of Biochemistry, introduction to Biomolecules, water as a biological solvent, weak acids and weak bases, pH and the concept of buffers, biological buffers and their physiological importance, the Henderson-Hasselbalch equation with simple numerical problems, and the common functional groups present in biomolecules.

MODULE 2: CARBOHYDRATES

15 Hours

Classification of carbohydrates; monosaccharides including their structures and Fischer and Haworth projections; reactions of monosaccharides with special reference to mutarotation; derivatives of monosaccharides such as amino sugars and glycosides; glycosidic bond formation along with disaccharides and oligosaccharides; polysaccharides including storage and structural polysaccharides; and bacterial cell wall polysaccharides.

MODULE 3: AMINO ACIDS & PROTEINS

15 Hours

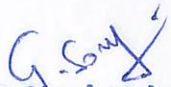
Classification, structure, stereochemistry, and chemical reactions of amino acids; the titration curve of glycine and pK values; essential, nonessential, and non-protein amino acids; peptide bond formation and naturally occurring peptides such as glutathione and enkephalin; protein classification and levels of structural organization including primary, secondary, tertiary, and quaternary structures with examples of haemoglobin and myoglobin; general properties of proteins including denaturation and renaturation; and determination of amino acid composition and amino acid sequencing of proteins.

MODULE 4: LIPIDS

15 Hours

Classification of lipids along with their reactions and properties; saturated, unsaturated, and essential fatty acids; the structure and functions of neutral fats, waxes, phospholipids, and sphingolipids; the structure and functions of cholesterol and glycolipids; prostaglandins and lipoproteins; bio membranes with emphasis on the behaviour of amphipathic lipids in water and the formation of micelles, bilayers, vesicles, and liposomes; and membrane composition with the fluid mosaic model.

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5. Reference Books

1. Dr. U.Satyanarayana and U.Chakrapani **Biochemistry** 5th ed. (2001)
Elsevier (New Delhi), Books and Allied Private Limited. ISBN: 81-87134-80-1
2. J.L.Jain : **Fundamentals of Biochemistry**, (2001), S. Chand & Company (New Delhi)
3. Albert L. Lehninger: **Principles of Biochemistry** (2013) 6th ed. Nelson, D.L. and Cox,
M.M.W.H. Freeman and Company (New York)
4. Jeremy M Berg, John L Tymoczko, and Lubert Stryer **Biochemistry**, 5th ed W H
Freeman; (2002) ISBN-10: 0-7167-3051-0
5. Dr. A.C Deb: **Fundamentals of Biochemistry**, (1999), New Central Book Agency
Private Limited. ISBN : 81-7381-144-X

6. Syllabus Focus

- a) Relevance to Local, Regional , National and Global Development Needs

Local/Regional/National /Global Development Needs	Relevance
Global Development needs	Understanding biomolecules has far-reaching implications, addressing global challenges and contributing to advancements in science and Diet therapies across the world.

- b) Components on Skill Development/Entrepreneurship Development/Employability

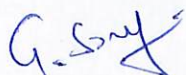
SD/ED/EMP	Syllabus Content	Description of Activity
SD	Module I, II, IV	Practicals- Qualitative analysis

7. Pedagogy

S. No	Type/Description of Activity	Student Centric Methods Adopted
1.	Science Experiments, Field Trips	Experiential Learning
2.	Presentation/Assignment	Participative Learning
3.	Quiz	Experiential Learning

8. Course Assessment Plana) **Weightage of Marks in Continuous Internal Assessments and End Semester Examination**

COs	Continuous Internal Assessments – CIA (40%)	End Semester Examination (60%)
CO1	CIA-1	End Semester examination
CO2	CIA-1	
CO3	CIA-2 – Objective	
CO4	CIA-2 – Assignment/ model making/ PPT	



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b) Model Question Paper

CHEMISTRY OF BIOMOLECULES

Code : U26/BIC/DSC/101
Credits : 4

Max Marks : 60 M
Time : 2 Hrs

I. Answer the following questions

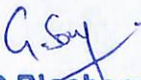
(4x10=40M)


1. Describe water as a biological solvent, highlighting its properties and biological significance.
OR
2. Define weak acids and weak bases. What is pH? Explain the concept of buffer solutions
3. Define carbohydrates. Give their classification with suitable examples indicating their structures.
OR
4. List the reactions of monosaccharides.
5. Explain the forces that stabilize the tertiary & quaternary structure of proteins.
OR
6. Describe the determination of N & C terminal amino acid of a peptide?
7. Illustrate the structural features & functions of phospholipids.
OR
8. Categorize the reactions of fatty acids & add a note on their biological importance.

II. Write Short notes on any 4 questions (out of 6)

(4x5=20M)

9. Henderson-Hasselbalch Equation
10. Bacterial cell wall Polysaccharides
11. Mutarotation
12. Denaturation of proteins
13. Essential amino acids
14. Lipoproteins


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Question Paper format – Blooms Taxonomy Level

SECTION A - INTERNAL CHOICE				4Q X 10 M = 40 M
Question Number	Module Covered	Question	CO	BTL (Blooms Taxonomy Level)
1	Module 1	Describe water as a biological solvent, highlighting its properties and biological significance.	CO 1	Level I
2	Module 1	Define weak acids and weak bases. What is pH? Explain the concept of buffer solutions	CO 1	Level I
3	Module 2	Define carbohydrates. Give their classification with suitable examples indicating their structures.	CO 2	Level I
4	Module 2	List the reactions of monosaccharides	CO 2	Level I
5	Module 3	Explain the forces that stabilize the tertiary & quaternary structure of proteins.	CO 3	Level II
6	Module 3	Describe the determination of N & C terminal amino acid of a peptide?	CO 3	Level II
7	Module 4	Illustrate the structural features & functions of phospholipids.	CO 4	Level III
8	Module 4	Categorize the reactions of fatty acids & add a note on their biological importance.	CO 4	Level III
SECTION B - ANSWER ANY 4 OUT OF 6 (To compulsorily have ONE question from each module)				4Q X 5 M = 20 M
9	Module 1	Henderson-Hasselbalch Equation	CO 1	Level I
10	Module 2	Bacterial cell wall Polysaccharides	CO 2	Level I
11	Module 2	Mutarotation	CO 2	Level I
12	Module 3	Denaturation of proteins	CO 3	Level II
13	Module 3	Essential amino acids	CO 3	Level II
14	Module 4	Lipoproteins	CO 4	Level III

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C. Question Paper Blueprint

Modules	Hours Allotted in the Syllabus	COs Addressed	Section A (No. of Questions)	Total Marks	Section B (No. of Questions)	Total Marks
1	15	CO-1	2	4X10=40	6 (Open Choice) At least 1 Question from each Module	4X5=20
2	15	CO-2	2			
3	15	CO-3	2			
4	15	CO-4	2			

9. CO-PO Mapping:

CO	PO	Cognitive Level	Class room sessions(hrs)
1	1	Remember	15
2	2	Understand	15
3	1	Application	15
4	2	Application	15


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**QUALITATIVE ANALYSIS OF BIOMOLECULES
PRACTICAL**

1. Course Description:

Max. Hours: 30

Course Code: U24/BIC/DSC/101/P

Type of course: DSC 1

Hours per week: 2

No. of credits: 1

Max. Marks: 50

2. Course objectives:

Introduce the basic molecules of life with respect to their isolations and qualitative estimations.

3. Course Outcomes:

This course will help the students to-

CO1: Demonstrate the procedures to isolate biomolecules from food sources.

CO2: Apply the skills in qualitative identification of Sugars and lipids by following a series of tests and procedures

4. Course Content

PRACTICAL SESSIONS

1. General Laboratory practices and safety procedures
2. Isolation of starch from potato.
3. Isolation of casein from milk.
4. Qualitative Analysis of Carbohydrates. (4 sessions)
5. Qualitative Analysis of Lipids. (3 sessions)
6. Preparation of buffers & determination of pH.
7. Achromic Point
8. Determination of specific rotation of sugars using Polarimeter (Glucose & Fructose)

5. Reference Books:

1. Experimental Biochemistry-A student companion -Beedu Sashidhar Rao & Vijay Deshpande
2. Laboratory Manual in Biochemistry-Jayaraman, J. Wiley Eastern

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6. Model Question Paper – End Semester Exam Practical

QUALITATIVE ANALYSIS OF BIOMOLECULES

Course Code: U26/BIC/DSC/101/P

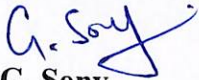

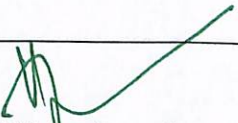
Credits: 1

Max Time: 2 Hrs


Max. Marks: 50


Answer the following.

1. Write the schematic representation (flowchart) for the qualitative analysis of
 - a. Carbohydrates
 - b. Principle involved in the isolation of casein from milk. (10 M)
2. Identify the given sugars present in the given solutions A and B (10 + 10 M)
3. Extract casein from the given sample of milk (10 M)
4. Viva (5M)
5. Record (5M)

Prepared by Course Teacher	Checked & verified by	Approved by
 Dr.G. Sony Teaching faculty	 Dr.G. Sony HOD Biochemistry	 Prof. Uma Joseph Principal

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SEMESTER - I

FUNDAMENTALS OF COMPUTERS

1. Course Description

Programme:	B. Sc	Max. Hours:	60
Course Code:	U26/CAP/DSC/101	Hours per week:	4
Type of Course:	DSC	Max. Marks:	60
No. of Credits:	4		

2. Course Objectives

1. To Study the Fundamentals of Computers and gain Knowledge of Operating systems.
2. To understand the significance of Microsoft Office tools.

3. Course Outcomes

After the successful completion of the course, the student will be able to:

CO1: Describe the evolution, types, and internal architecture of computer systems, types of operating systems, perform basic DOS and GUI operations, and identify security threats like viruses and hacking. (L IV)

CO2: Understand, edit and manage professional documents using word processing software and design and deliver engaging presentations using multimedia elements, animations. (L II)

CO3: To learn the skills for working with formulas, functions, named ranges, referencing cells and auditing for effective data analysis and manipulation. (L II)

CO4: Demonstrate ability to employ analytical tools and visualization techniques to extract insights and present findings effectively. (L III)

4. Course content

MODULE I:

15 Hours


INTRODUCTION TO INFORMATION TECHNOLOGY: Development of computers- Generations of computers- An overview of computer system- Types of computers- Input & Output Devices. Looking inside the machine: Basic components of a computer system- Control unit- ALU- Input/output functions- Memory- RAM- ROM- EPROM- PROM and Other types of memory. **Operating System:** Meaning- Definition & Functions- Types of OS- Booting process- DOS- Commands (internal & external)- GUI- wild card characters- Virus & Hackers- Cryptography & cryptology. Windows: Using the Start Menu- Control Panel- Using multiple windows- Customizing the Desktop Windows accessories (Preferably latest version of Windows or Linux Ubuntu)


MODULE II:

15 Hours

WORD PROCESSING: Application of word processing software- Menus & Tool Bars- Opening word processor- Creating- Entering- Saving & printing the document- Editing &

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Formatting, Tables, Text Mail Merge and Macros (Preferably latest version of MS Word or Libre Office Writer).

PRESENTATION: Application of Presentation– Menus & Tool bars– Creating presentations– Adding- Editing and deleting slides- Templates and manually– Slide show– Saving- Opening and closing a Presentation–Types of slides- Slide Views- Formatting– Insertion of Objects and Charts in slides- Custom Animation and Transition.

MODULE III:

15 Hours

DATA REPRESENTATION: Review of Excel-Worksheet Basics, Protecting Workbook Importing and Exporting data, sharing in Excel; Formula sand functions – understanding formulas; operators in formula; named ranges; calculations; functions in formulas; relative and absolute addressing; referencing cells outside the worksheet and workbook; functions- logical, summarizing, text, lookup, reference, data and time, math functions; error handling, formula auditing.


MODULE IV:

15 Hours

DATA VISUALIZATION: Charts–waterfall, histogram, pareto, box and whisker, Tree map, sunburst, sparkline,3D Map charts and their uses; Advanced charts -Milestone chart, SmartArt graphics, Organization chart. Pivot tables–verify data source, format data, recommended pivot tables, adding slicers, timelines, calculated fields and group fields in pivot.

5. Reference Books

1. **Fundamentals of Computers**, Reema Thareja, 2nd Edition 2019. (Module 1)
2. **Microsoft Office Step by Step (Office 2021 and Microsoft 365)**, Joan Lambert, 1st edition, 2022 (Module 2)
3. **Computer Basics with Office Automation**, Archana Kumar, Wiley publications, 2019. (Module 2)
4. **Fundamentals of Information Technology**: Dr. NVN Chary, Kalyani Publishers. (Module 3)
5. **A text book of Information Technology**: Rachhpal Singh, Kalyani Publishers (Module 3.)
6. **“Data Analysis with Excel”**, Manisha Nigam, BPP publications, 2018 (Module 3 &4)


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6. Syllabus Focus

a) Relevance to Local, Regional, National and Global Development Needs

S. No	Student Centric Methods Adopted	Type/Description of Activity
1.	Global Development	Data Analysis is a systematic method to look for trends, groupings, or other relationships between different types of data. Data visualization is to make it easier to identify patterns, trends and outliers in large data sets.
2.	Global Development	Basic computer skills such as word processing, spreadsheets, presentations, and the internet, are essential for most jobs and are considered valuable skills in the workforce. Good computer skill aligns with an individual's career goals and enhances productivity and effectiveness in the workplace.

b) Components on Skill Development/Entrepreneurship Development/Employability

SD/ED/EMP	Syllabus Content	Description of Activity
Skill Development	Module III & IV	Hands- on Practicals



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7. Pedagogy

S. No	Student Centric Methods Adopted	Type / Description of Activity
1.	Hands-on Practical Sessions	Students perform step-by-step exercises in computer labs to practice basic computer operations, spreadsheet usage, and data visualization tools.
2.	Problem-Solving Activities	Students are given practical data problems (sorting, filtering, creating charts) and asked to find solutions using computer tools.
3.	Collaborative Learning	Students work in groups to collect, clean, and analyze data, and present their findings using graphs and visualizations.

8. Course Assessment Plan

a) Weightage of Marks in Continuous Internal Assessments and End Semester Examination

CO	Continuous Internal Assessments CIA - 40%	End Semester Examination- 60%
CO1	CIA 2 – Test 1: MCQ's, Quiz test or subjective	Written Exam
CO2	CIA 1 - Subjective	
CO3		
CO4	CIA 2 – Test 2: MCQ's or Presentation	


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Dr. Sujatha Yemina

b) Model Question Paper – End Semester Exam Theory

FUNDAMENTALS OF COMPUTERS

Course Code: U26/ CAP/ DSC/101

Credits: 4

MAX MARKS: 60

TIME: 2 hours

Note: This question paper consists of Section A and B. The answer to Section A & B must be written in the answer book given.

SECTION – A (Long Essay Type)

Answer ALL questions:

Marks: 4 x 10 =40

1. Explain the basic components of a computer system.

OR

2. Explain the operating system in detail.

3. Explain text creation and formatting in MS-Word.

OR

4. Explain the features used for creating presentations in PowerPoint.

5. Explain different types of Excel functions.

OR

6. Explain the basics of Excel worksheets.

7. Explain the use of slicers, timelines, calculated fields, and group fields in Pivot Tables.

OR

8. Explain Pivot Tables in Excel.

SECTION –B (Short Essay Type)

Write short notes on any FOUR of the following:

Marks: 4 x 5=20

9. Write short notes on:

a) Wild card characters

b) Computer viruses and hackers

10. Explain GUI (Graphical User Interface)

11. Write short notes on Tree Map and Sunburst charts.

12. Explain error handling and formula auditing in Excel.

13. What is cell referencing?

14. What are slide transitions and animations in PowerPoint?

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Question Paper format – Blooms Taxonomy Level

SECTION A - INTERNAL CHOICE				4Q X 10 M = 40 M
Question Number	Module Covered	Question	CO	BTL (Blooms Taxonomy Level)
1	Module 1	Explain the basic components of a computer system.	CO 1	Level IV
2	Module 1	Explain the operating system in detail.	CO 1	Level IV
3	Module 2	Explain text creation and formatting in MS-Word.	CO 2	Level II
4	Module 2	Explain the features used for creating presentations in PowerPoint.	CO 2	Level II
5	Module 3	Explain different types of Excel functions.	CO 3	Level II
6	Module 3	Explain the basics of Excel worksheets.	CO 3	Level II
7	Module 4	Explain the use of slicers, timelines, calculated fields, and group fields in Pivot Tables.	CO 4	Level III
8	Module 4	Explain Pivot Tables in Excel.	CO 4	Level III
SECTION B - ANSWER ANY 4 OUT OF 6 (To compulsorily have ONE question from each module)				4Q X 5 M = 20 M
9	Module 1	Write short notes on: a) Wild card characters b) Computer viruses and hackers	CO 1	Level IV
10	Module 1	Explain GUI (Graphical User Interface)	CO 1	Level IV
11	Module 2	What are slide transitions and animations in PowerPoint?	CO 2	Level II
12	Module 3	What is cell referencing?	CO 2	Level II
13	Module 3	Explain error handling and formula auditing in Excel.	CO 3	Level II
14	Module 4	Write short notes on Tree Map and Sunburst charts.	CO 4	Level III

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c) Question Paper Blueprint

Modules	Hours Allotted in the Syllabus	CO Addressed	Section A (No. of Questions)	Total Marks	Section B (No. of Questions)	Total Marks
1	15	CO-1	2	4x10=40	6	4x5=20
2	15	CO-2	2		(By taking at least one question from each Module)	
3	15	CO-3	2			
4	15	CO-4	2			

9. CO-PO Mapping

CO	PO	Cognitive Level	Classroom sessions (hrs)
1	1, 2	Analyze	15
2	2, 7	Understand	15
3	1, 2	Remember	15
4	1, 2	Understand	15


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FUNDAMENTALS OF COMPUTERS

Practical Syllabus

1. Course Description

Programme:	B. Sc	Max. Hours:	30
Course Code:	U26/CAP/DSC/101/P	Hours per week:	2
Type of Course:	DSC	Max. Marks:	50
No. of Credits:	1		

2. Course Objectives

1. To understand the basics of computer systems, their developments, generations of Computers
2. To provide knowledge of operating systems, their types, functions, booting processes, and popular environments such as Windows and Linux.

3. Course Outcomes

After the successful completion of the course, the student will be able to:


CO1: Interpret basics of computers and **Use** word processing software (Cognitive levels – 3)


CO2: Use Spreadsheets and Presentation Software (Cognitive level – 3)

4. Course Content

1. Demonstrate basic DOS commands like DIR, COPY, DEL, MD, CD, TYPE, DATE, TIME, and differentiate between internal & external commands.
2. Create a formal letter/document using formatting tools (font styles, sizes, line spacing, bullets, numbering).
3. Use Mail Merge to send an invitation to a list of people using a sample address list.
4. Create and run a Macro that automates formatting (e.g., bold, font size 14, centre-align).
5. Insert and format images, tables, and hyperlinks in a document.
6. Demonstrate use of Headers, Footers, Page Numbering and print preview before printing a document.
7. Create a mark sheet using a spreadsheet and calculate total, average, grade using formulas and conditional formatting.
8. Apply sorting and filtering on a sample data of students with name, marks, and grade.
9. Exercises using Google Docs
10. Exercises in Google Sheets

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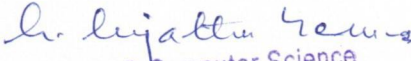
11. Create charts (bar, line, pie) to visually represent sales or performance data.
12. Create a 5-slide presentation on any topic using different slide layouts and themes.
13. Insert multimedia elements such as images, videos, and audio clips into slides.
14. Apply custom animations and transitions to slides and run the slideshow.
15. Use charts and Smart Art to represent data in a visually engaging format.
16. Demonstrate how to send an email with attachments using Gmail or any mail client.
17. Download and insert multimedia (image/audio/video) in a Word or PowerPoint file.
18. Refer the excel worksheet shared and perform the below tasks:
 - a. Apply conditional Formatting for the Total column and find how many students Mark1 is greater than 42.
 - b. Calculate total and Percentage.
 - c. Use the VLOOKUP function to find the name of the student with Roll No 12823.
 - d. Use IF CONDITION for Pass/fail. If Percentage greater than 72 "Pass" else "Fail".
 - e. Add a new column as Full name and fill the data using the CONCATENATE function.
19. Draw a sparkline chart for the following chart

Region	Q1	Q2	Q3	Q4
Asia	\$ 140,787	\$143,549	\$46,210	\$66,404
Europe	\$78,944	\$115,252	\$198,783	\$182,729
South America	\$100,961	\$119,646	\$83,233	\$99,524
North America	\$82,889	\$68,945	\$187,202	\$110,372

5. Reference Books

1. Microsoft Office Step by Step (Office 2021 and Microsoft 365), Joan Lambert, 1st edition, 2022
2. Computer Basics with Office Automation, Archana Kumar, Wiley publications, 2019
3. Introduction to Computers, Peter Norton, McGraw-Hill, 2012
4. Manisha Nigam, "Data Analysis with Excel", BPP publications, 2018


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Model Question Paper – End Semester Exam Practical

FUNDAMENTALS OF COMPUTERS

Programme: B.Sc.

Course Code: U26/CAP/DSC/101/P

Type of Course: DSC

No. of credits: 1

Max. Marks: 50

Section A

Answer the circled question

1. Use Mail Merge to send an invitation to a list of people using a sample address list.
2. Create a 5-slide presentation on any topic using different slide layouts and themes.

Section B

Answer any one

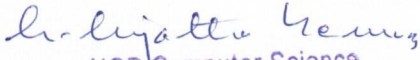
1. Draw a sparkline chart for the following chart.


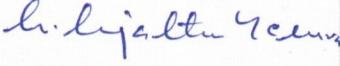

Region	Q1	Q2	Q3	Q4
Asia	\$ 140,787	\$143,549	\$46,210	\$66,404
Europe	\$78,944	\$115,252	\$198,783	\$182,729
South America	\$100,961	\$119,646	\$83,233	\$99,524
North America	\$82,889	\$68,945	\$187,202	\$110,372

4. Refer the excel worksheet shared and perform the below tasks:

- a. Apply conditional Formatting for the Total column and find how many students Mark1 is greater than 42.
- b. Calculate total and Percentage.
- c. Use the VLOOKUP function to find the name of the student with Roll No 12823.
- d. Use IF CONDITION for Pass/fail. If Percentage greater than 72 "Pass" else "Fail".
- e. Add a new column as Full name and fill the data using the CONCATENATE Function.


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